Naming Ionic Compounds Worksheet Answer Key PDF

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Part 1: Building a Foundation

What is the suffix used for naming anions in binary ionic compounds?

undefined. -ate

undefined. -ide ✓

undefined. -ite

undefined. -ous

The suffix used for naming anions in binary ionic compounds is '-ide'.

Which of the following are characteristics of ionic compounds? (Select all that apply)

undefined. A) Composed of metals and non-metals ✓

undefined. B) Have high melting points ✓

undefined. C) Conduct electricity in solid state

undefined. D) Form crystal lattice structures ✓

Ionic compounds are composed of metals and non-metals, have high melting points, conduct electricity in solid state, and form crystal lattice structures.

Explain why ionic compounds are electrically neutral.

Ionic compounds are electrically neutral because the total positive charge from cations equals the total negative charge from anions.

List two examples of polyatomic ions and their charges.

1. Example 1

Sulfate (SO4^2-)

2. Example 2

Nitrate (NO3^-)

Examples include sulfate (SO4^2-) and nitrate (NO3^-).

Part 2: Comprehension and Application

Which of the following correctly describes the naming of a compound with a transition metal?

undefined. A) The metal is named first with its charge in Roman numerals. ✓
undefined. B) The non-metal is named first with its charge in Roman numerals.
undefined. C) The metal is named with the suffix "-ide."
undefined. D) The non-metal is named with the suffix "-ate."

The correct description is that the metal is named first with its charge in Roman numerals.

When naming ionic compounds, which of the following statements are true? (Select all that apply)

undefined. A) The cation is always named first. ✓
undefined. B) Anions are named using the suffix "-ate."
undefined. C) The formula must reflect a neutral charge. ✓
undefined. D) Transition metals do not require charge specification.

The cation is always named first, the formula must reflect a neutral charge, and transition metals require charge specification.

Describe the process of naming an ionic compound containing a polyatomic ion.

The process involves identifying the cation and anionic part, naming the cation first, and then naming the polyatomic ion.

What is the correct name for the compound with the formula Na2SO4?

undefined. A) Sodium Sulfide
undefined. B) Sodium Sulfate ✓
undefined. C) Sodium Sulfite
undefined. D) Sodium Sulfate(IV)

The correct name for Na2SO4 is Sodium Sulfate.

Which of the following formulas correctly represent ionic compounds? (Select all that apply)

undefined. A) CaCl2 ✓
undefined. B) Na2O ✓
undefined. C) Mg2S ✓
undefined. D) Al2O3 ✓

The correct formulas representing ionic compounds are CaCl2, Na2O, Mg2S, and Al2O3.

Write the chemical formula for the compound formed between calcium ions and nitrate ions.

The chemical formula for the compound formed is Ca(NO3)2.

Part 3: Analysis, Evaluation, and Creation

If a compound is formed between Fe3+ and Cl-, what is the correct formula?

undefined. A) FeCl
undefined. B) FeCl2
undefined. C) FeCl3 ✓
undefined. D) Fe3Cl

The correct formula for the compound is FeCl3.

Analyze the following statements and select those that are true about ionic compound formation. (Select all that apply)

undefined. A) The total positive charge must equal the total negative charge. ✓
undefined. B) Ionic compounds can have a net charge.
undefined. C) Ionic compounds are typically soluble in water. ✓
undefined. D) The formula of an ionic compound reflects the ratio of ions. ✓

The true statements are that the total positive charge must equal the total negative charge, ionic compounds are typically soluble in water, and the formula reflects the ratio of ions.

Explain how the charge of a transition metal affects the naming and formula of an ionic compound.

The charge of a transition metal is indicated by Roman numerals in the name, which helps determine the formula of the ionic compound.

Which of the following scenarios best illustrates the importance of charge balance in ionic compounds?

undefined. A) Mixing two metals to form an alloy.

undefined. B) Dissolving salt in water and observing conductivity.

undefined. C) Creating a model of a crystal lattice.

undefined. D) ObservING the reaction between sodium and chlorine gas. ✓

The scenario that best illustrates the importance of charge balance is observing the reaction between sodium and chlorine gas.

Evaluate the following compounds and determine which are named correctly. (Select all that apply)

undefined. A) KCl - Potassium Chloride ✓

undefined. B) FeO - Iron(II) Oxide ✓

undefined. C) CuSO4 - Copper(II) Sulfate ✓

undefined. D) Al(NO3)3 - Aluminum Nitrate ✓

The correctly named compounds are KCl - Potassium Chloride, FeO - Iron(II) Oxide, CuSO4 - Copper(II) Sulfate, and Al(NO3)3 - Aluminum Nitrate.

Propose a method for teaching the naming of ionic compounds to students who are new to chemistry. Include at least two teaching strategies.

One method could be using visual aids to illustrate ionic bonds, and another could be incorporating interactive activities to practice naming.