

Periodic Trends Quiz PDF

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What is the trend in metallic character as you move down a group in the periodic table?

- Increases
- Decreases
- Remains constant
- Varies randomly

Which of the following statements about ionization energy are true? (Select all that apply)

- It decreases down a group.
- It increases across a period.
- It is the energy required to add an electron.
- It is the energy required to remove an electron.

Which of the following factors affect atomic radius? (Select all that apply)

- Number of electron shells
- Nuclear charge
- Ionization energy
- Electron affinity

What is the general trend of ionization energy across a period in the periodic table?

- Decreases
- Increases
- Remains constant
- Fluctuates

Which of the following elements has the largest atomic radius?

- Fluorine
- Oxygen

- Nitrogen
- Carbon

What are characteristics of nonmetals? (Select all that apply)

- High electronegativity
- Good conductors of electricity
- brittle in solid form
- High ionization energies

Which of the following elements is a metalloid?

- Silicon
- Sulfur
- Phosphorus
- Argon

In which group of the periodic table are the alkali metals found?

- Group 1
- Group 2
- Group 17
- Group 18

Which element is the most electronegative?

- Oxygen
- Chlorine
- Fluorine
- Nitrogen

Which of the following elements have high reactivity? (Select all that apply)

- Fluorine
- Helium
- Sodium
- Neon

Which of the following trends are observed across a period from left to right? (Select all that apply)

- Increase in ionization energy
- Decrease in atomic radius
- Increase in metallic character
- Increase in electronegativity

Which elements are considered transition metals? (Select all that apply)

- Iron
- Copper
- Sodium
- Zinc

Which of the following elements has the highest electron affinity?

- Sodium
- Chlorine
- Argon
- Lithium

Which of the following best describes the periodic law?

- Elements are arranged by increasing atomic mass.
- Elements are arranged by increasing atomic number.
- Elements are arranged by decreasing atomic number.
- Elements are arranged by decreasing atomic mass.