

Periodic Trends Quiz Answer Key PDF

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What is the trend in metallic character as you move down a group in the periodic table?

- A. Increases ✓**
- B. Decreases
- C. Remains constant
- D. Varies randomly

Which of the following statements about ionization energy are true? (Select all that apply)

- A. It decreases down a group. ✓**
- B. It increases across a period. ✓**
- C. It is the energy required to add an electron.
- D. It is the energy required to remove an electron. ✓**

Which of the following factors affect atomic radius? (Select all that apply)

- A. Number of electron shells ✓**
- B. Nuclear charge ✓**
- C. Ionization energy
- D. Electron affinity

What is the general trend of ionization energy across a period in the periodic table?

- A. Decreases
- B. Increases ✓**
- C. Remains constant
- D. Fluctuates

Which of the following elements has the largest atomic radius?

- A. Fluorine
- B. Oxygen
- C. Nitrogen
- D. Carbon ✓**

What are characteristics of nonmetals? (Select all that apply)

- A. High electronegativity ✓**
- B. Good conductors of electricity
- C. brittle in solid form ✓**
- D. High ionization energies ✓**

Which of the following elements is a metalloid?

- A. Silicon ✓**
- B. Sulfur
- C. Phosphorus
- D. Argon

In which group of the periodic table are the alkali metals found?

- A. Group 1 ✓**
- B. Group 2
- C. Group 17
- D. Group 18

Which element is the most electronegative?

- A. Oxygen
- B. Chlorine
- C. Fluorine ✓**
- D. Nitrogen

Which of the following elements have high reactivity? (Select all that apply)

- A. Fluorine ✓**
- B. Helium

C. Sodium ✓

D. Neon

Which of the following trends are observed across a period from left to right? (Select all that apply)

A. Increase in ionization energy ✓

B. Decrease in atomic radius ✓

C. Increase in metallic character

D. Increase in electronegativity ✓

Which elements are considered transition metals? (Select all that apply)

A. Iron ✓

B. Copper ✓

C. Sodium

D. Zinc ✓

Which of the following elements has the highest electron affinity?

A. Sodium

B. Chlorine ✓

C. Argon

D. Lithium

Which of the following best describes the periodic law?

A. Elements are arranged by increasing atomic mass.

B. Elements are arranged by increasing atomic number. ✓

C. Elements are arranged by decreasing atomic number.

D. Elements are arranged by decreasing atomic mass.