

Exothermic Reactions Quiz PDF

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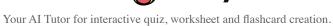
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What is the sign of the enthalpy change (ΔH) for an exothermic reaction?		
○ Positive		
○ Negative		
Zero		
Undefined		
Which of the following is an example of an exothermic reaction?		
○ Melting ice		
○ Photosynthesis		
CombustION of propane		
○ Electrolysis of water		
Which of the following best describes an exothermic reaction?		
Absorbs heat from the surroundings		
Releases heat to the surroundings		
Has no change in temperature		
Requires constant heating to proceed		
What happens to the temperature of the surroundings during an exothermic reaction?		
○ It decreases		
O It remains constant		
O It increases		
O It fluctuates		
What is typically required to initiate an exothermic reaction?		
○ Catalyst		
○ Activation energy		



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Endothermic reactionEquilibrium state
Which of the following reactions is exothermic?
 Boiling water Neutralization of HCl and NaOH Sublimation of dry ice Decomposition of water
In an exothermic reaction, the energy of the products is:
 Higher than the reactants Lower than the reactants Equal to the reactants Unrelated to the reactants
In an exothermic reaction, which of the following statements are true?
 □ Products have higher energy than reactants □ Heat is absorbed from the surroundings □ Heat is released to the surroundings □ Products have lower energy than reactants
Which processes involve exothermic reactions?
 □ Formation of rust □ Dissolving sugar in water □ BurnING wood □ Cooking an egg
Explain why the combustion of fossil fuels is considered an exothermic reaction.





Describe how calorimetry can be used to measure the heat released in an exothermic reaction.		
Discuss the environmental impact of exothermic reactions, particularly in the context of fossil fuel combustion.		
Explain the role of activation energy in exothermic reactions and how it affects the reaction rate.		
Describe a real-world application of an exothermic reaction and explain its significance.		



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Explain how exothermic reactions contribute to the spontaneity of a chemical process.
Which process is exothermic?
Evaporation of water
O Dissolving ammonium nitrate in water
Freezing of water
○ Photosynthesis
Which of the following are examples of exothermic reactions?
☐ CombustION of gasoline
Photosynthesis
☐ Respiration in cells☐ Melting of ice
Metting of ice
Which of the following safety considerations are important for exothermic reactions?
☐ Proper ventilation
Use of fire retardants
Monitoring temperature
☐ Increasing pressure
What are the implications of exothermic reactions in industrial applications?
☐ Energy efficiency
☐ Increased pollution
☐ Cost savings
Requirement for cooling systems
Which of the following are characteristics of exothermic reactions?



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☐ Energy is absorbed	
\square $\triangle H$ is negative	
☐ Energy is released	
☐ Temperature of surroundings decreases	
☐ Temperature of surroundings decreases	