

Endothermic Reactions Quiz PDF

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What happens to the temperature of the surroundings during an endothermic reaction?

- Increases
- Decreases
- Stays the same
- Fluctuates

In an endothermic reaction, how does the potential energy of the products compare to the reactants?

- Lower
- Higher
- Equal
- Can't be determined

Which of the following reactions is endothermic?

- Combustions of wood
- Photosynthesis
- Rust of iron
- Neutralization of acid and base

Which of the following is a real-life application of an endothermic reaction?

- Hand warmers
- Cold packs
- Fireworks
- Incandescent bulbs

What is the sign of the enthalpy change (ΔH) for an endothermic reaction?

- Negative

- Zero
- Positive
- Variable

What is required to initiate most endothermic reactions?

- Catalyst
- Light
- Activation energy
- Water

Which of the following best describes an endothermic reaction?

- Releases heat
- Absorbs heat
- Produces light
- Emits sound

Provide a detailed example of a real-life application of an endothermic reaction and explain how it works.

Compare and contrast endothermic and exothermic reactions in terms of energy flow and enthalpy change.

Which of the following industrial processes involve endothermic reactions? (Select all that apply)

- Haber process
- Electrolysis of water
- Fermentation
- Production of quicklime

Which of the following are characteristics of endothermic reactions? (Select all that apply)

- Absorb heat
- Release heat
- Have a positive ΔH
- Decrease the temperature of surroundings

Explain why photosynthesis is considered an endothermic reaction.

Which processes are considered endothermic? (Select all that apply)

- Boiling water
- Freezing water
- Photosynthesis
- Combustions of coal

Describe the role of activation energy in endothermic reactions.

In an energy diagram for an endothermic reaction, which of the following are true? (Select all that apply)

- Reactants have higher energy than products
- Products have higher energy than reactants
- Energy is absorbed
- Energy is released

Which of the following statements about endothermic reactions are correct? (Select all that apply)

- They are always spontaneous
- They require energy input to proceed
- They result in a temperature decrease in the surroundings
- They have a negative enthalpy change

Which of the following is an example of an endothermic process?

- Combustions of gasoline
- Freezing of water
- Melting of ice
- Condensation of steam

Discuss the environmental impacts of endothermic reactions compared to exothermic reactions.

What are common signs of an endothermic reaction occurring? (Select all that apply)

- Temperature increase
- Temperature decrease
- Absorption of heat
- Light emission

How can calorimetry be used to measure the heat absorbed in an endothermic reaction?

